



# 10TH SSC MCQ - CH - PERIODIC CLASSIFICATION OF ELEMENT

DATE:

TIME: 31 Min

MARKS: 31

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**Note:-**

1. All Questions are compulsory.
2. Numbers on the right indicate full marks.

- Q.1 The elements of the group 16 are also called (1)  
A) halogens      B) noble gases  
C) chalcogens    D) alkaline earth metals

**Ans : C**

- Q.2 Most metallic element in the fifth period is (1)  
A) silver          B) rubidium  
C) gold            D) rhodium

**Ans : B**

- Q.3 In the periodic table from left to right in a period the atomic volume (1)  
A) decreases      B) increases  
C) remains same    D) first decreases then increases

**Ans : D**

- Q.4 Which of the following remains unchanged on descending a group in the periodic table? (1)  
A) Valence electrons    B) Atomic size  
C) Density                D) Metallic character

**Ans : A**

- Q.5 The amount of energy released on the addition of an electron in outermost shell of an atom (1)  
1s called  
A) ionization energy    B) hydration energy  
C) electron affinity      D) electron negativity

**Ans : C**

- Q.6 Electron affinity depend on (1)  
A) Atomic size          B) Nuclear charge  
C) Atomic number      D) Both atomic size and nuclear charge

**Ans : D**

- Q.7 The number of elements present in fifth period is (1)  
A) 18                  B) 6  
C) 8                    D) 32

**Ans : A**

- Q.8 Gradual addition of electronic shell in the noble gases causes a decrease in their (1)  
A) Ionization energy    B) Atomic radius  
C) Boiling point        D) Density

**Ans : A**

- Q.9 Heaviest atom among the following is (1)  
A) uranium      B) lead  
C) mercury      D) radium

**Ans : A**

- Q.10 Variable valency in general is shown by (1)  
A) gaseous elements      B) non-metals  
C) s-block elements      D) d-block elements

**Ans : D**

- Q.11 When first ionization energy is plotted against the atomic number, then peaks in curve are occupied by (1)  
A) halogens      B) rare gases  
C) alkali metals      D) transition elements

**Ans : B**

- Q.12 The elements of group 1 are called alkali metals because (1)  
A) the metals are corroded by alkali      B) their oxides are alkaline  
C) their hydrides are strongly alkaline      D) none of these

**Ans : B**

- Q.13 Ionization potential is (1)  
A) The tendency of an atom to lose an electron      B) The tendency of an anion to lose electron  
C) The tendency of cation to gain electron      D) The ratio of the cation to anion

**Ans : A**

- Q.14 Lanthanides and actinides are recognised as (1)  
A) s- block element      B) d-block element  
C) f-block element      D) p-block element

**Ans : C**

- Q.15 The most non-metallic element is (1)  
A) Sodium      B) Helium  
C) Chlorine      D) Fluorine

**Ans : D**

- Q.16 The lightest metal in periodic table is (1)  
A) Magnesium      B) Zinc  
C) Lithium      D) Sodium

**Ans : C**

- Q.17 Which of the following has highest electronegativity? (1)  
A) Carbon      B) Magnesium  
C) Oxygen      D) Sulphur

**Ans : C**

- Q.18 The elements with atomic numbers 4, 12, 20, 38, 56, 88 belong to which of the given group? (1)  
A) Alkali metals    B) Alkaline earth metals  
C) Inert gases      D) Halogens

**Ans : B**

- Q.19 The lanthanides are placed in the periodic table at the \_\_\_\_\_ (1)  
A) Left hand side    B) right hand side  
C) middle            D) bottom

**Ans : D**

- Q.20 Halogens are placed in which group of elements in modern periodic table? (1)  
A) 17                B) 2  
C) 4                 D) 6

**Ans : A**

- Q.21 The correct formula of the oxide of eka-aluminium element predicted by Mendeleev was \_\_\_\_\_ (1)  
A)  $EO_3$             B)  $E_3O_2$   
C)  $E_2O_3$             D) (d)  $EO$

**Ans : C**

- Q.22 Which of the following element would lose an electron easily? (1)  
A) Mg                B) Ca  
C) Na                D) K

**Ans : D**

- Q.23 The atomic number of an element is 20. In modern periodic table, this element is placed in \_\_\_\_\_ (1)  
A) 2<sup>nd</sup> period    B) 4<sup>th</sup> period  
C) 3<sup>rd</sup> period    D) 1<sup>st</sup> period

**Ans : B**

- Q.24 The modern periodic table was prepared by \_\_\_\_\_ (1)  
A) Newlands    B) Dalton  
C) Moseley      D) Mendeleev

**Ans : C**

- Q.25 \_\_\_\_\_ of the following triads does not follow Dobereiner's law of triads. (1)  
A) Li, Na, K      B) Ca, Sr, Ba  
C) Cl, Br, I        D) Cu, Ag, Au

**Ans : D**

- Q.26 Which of the following sets belong to the same period? (1)  
A) Li, Na, K      B) Li, Mg, Ca  
C) Ni, Cu, Zn    D) F, Cl, Br

**Ans : C**

Q.27 According to Newlands octave, chlorine shows similarity with \_\_\_\_\_ (1)

- A) Beryllium    B) Fluorine  
C) lithium      D) Sodium

**Ans : B**

Q.28 In which block of the modern periodic table are the non-metals found? (1)

- A) s-block      B) p-block  
C) d-block      D) f-block

**Ans : B**

Q.29 Molecular formula of the chloride of an element X is XCl. This compound is a solid having high melting point. Which of the following elements be present in the same group as X ? (1)

- A) Na            B) Mg  
C) Al            D) Si

**Ans : A**

Q.30 Alkaline earth metals have valency 2. This means that their position in the modern periodic table is in \_\_\_\_\_ (1)

- A) Group 2      B) Group 16  
C) Period 2      D) d-block

**Ans : B**

Q.31 The number of electrons in the outermost shell of alkali metals is \_\_\_\_\_ (1)

- A) 1              B) 2  
C) 3              D) 7

**Ans : A**